

T09D02 – (9.3) Reactivity

Name.....

1. 9.3.1 Deduce a reactivity series based on the chemical behavior of a group of oxidizing and reducing agents. (3)
- a. Why do metals generally behave as reducing agents?

- b. Provide an example of a replacement reaction where Mg takes the place of Cu^{2+} in the compound CuSO_4 :

i. Write the Total Ionic Equation:

ii. Write the Net Ionic Equation:

iii. Provide each of the $\frac{1}{2}$ reactions:

iv. Which metal is more reactive?

- c. Describe and diagram a thermite reaction:

- d. What happens when common metals are added to water? Use Na and H_2O as an example:

2. 9.3.2 Deduce the feasibility of a redox reaction from a given reactivity series. (3)